



NAMING COMPOUNDS OF METALS THAT FORM TWO KINDS OF POSITIVE IONS

IONIC COMPOUND:

METAL + NON METAL

BLUE + YELLOW
(Colors from periodic table)

ELEMENTAL + ELEMENTAL
NAME NAME
PLUS: ROMAN PLUS: IDE
NUMBERS

NO PREFIXES

COVALENT COMPOUND:

NON METAL + NON METAL

YELLOW + YELLOW
(Colors from periodic table)

ELEMENTAL + ELEMENTAL
NAME NAME
PLUS: IDE

PREFIXES

↓

HOW TO NAME AN IONIC COMPOUND OF METALS THAT FORM TWO KINDS OF POSITIVE IONS:

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METALS WITH TWO KIND OF POSITIVE IONS:

(Remember these elements)

Au^+ (Gold I)	Cu^+ (Copper I)	Fe^{2+} (Iron II)	Sn^{2+} (Tin II)	Pb^{2+} (Lead II)
Au^{3+} (Gold III)	Cu^{2+} (Copper II)	Fe^{3+} (Iron III)	Sn^{4+} (Tin IV)	Pb^{4+} (Lead IV)

↓

EXAMPLES: IRON (II) CHLORIDE (FeCl_2)

FeCl_2 → Chlorine's subscript 2 comes from Iron → $\text{Fe}^2 \text{Cl}^1$
Iron's number 1 (invisible) comes from Chlorine
because Fe behaves as Iron II

COPPER (I) OXIDE (Cu_2O)

Cu_2O → Copper's subscript 2 comes from Oxygen → $\text{Cu}^1 \text{O}^2$
Oxygen's number 1 (invisible) comes from Copper
because Cu behaves as Copper I